**1.2.2 EXERCISE 4 - ELECTRONEGATIVITY AND BOND POLARITY**

1. Explain the meaning of the term electronegativity.

2. Draw the correct dipoles on the following covalent bonds:

 a) Cl – Cl

 b) H – O

 c) B – F

 d) N – F

 e) N – O

 f) H – Cl

 g) H – N

 h) B – H

3. Draw the following molecules, add dipoles to all the bonds and predict whether or not they have an overall dipole:

 a) CO2

 b) SO2

 c) BF3

 d) NF3

 e) Cl2

 f) HCl