CHEM 121 2018 CA ASSIGNMENT MARK SCHEME

Note: extra sf will not be penalised but rounding errors which result in an answer not correct to 3sf will be penalised

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| **1.** | (a) | C8H18 + 12.5O2 🡪 8CO2 + 9H2O (1) |
|  | (b) | moles of octane = 1000/114 = 8.77 moles (1)energy released = 5740 x 8.77 = 50,400 kJ (1) |
|  | (c) | Moles of octane = 100000/5740 = 17.4 moles so moles of CO2 = 17.4 x 8 = 139 moles (1)Mass of CO2 = 139 x 44 = 6132 g or 6.13 kg (1)[5] |
| **2.** | (a) | ΔT = 45.4 – 18.1 = 27.3 oC and mass of alcohol burned = 208.8 – 208.58 = 0.22 g (both = 1)q = 50 x 4.18 x 27.3 = 5706 J (1)moles of alcohol = 0.22/74 = 0.00297 (1)molar enthalpy change = 5706/0.00297 = -1920 kJmol-1 (1) must be negative |
|  | (b) | Heat loss from the flame to the surrounding air (1)[5] |
| **3.** | (a) | 2AgNO3 + Zn 🡪 Zn(NO3)2 + 2Ag (1) |
|  | (b) | q = 3.2 x 4.18 x 50 = 668.8 J (1) |
|  | (c) | Moles of AgNO3 = 0.2 x 50/1000 = 0.01 (1)So moles of Zn = 0.01/2 = 0.005 (1)Molar enthalpy change = 668.8/0.005 = -133760 Jmol-1 or -134 kJmol-1 must be negative (1)[5] |
| **4.** | (a) | Bonds broken = 4 C-H, 1 C-C, (1 C=O), 2.5 O=O = 3980 (or 3244) (1)Bonds formed = 4 x O-H, 3/4 x C=O = 4800 (or 4064) (1)ΔH = 3980 - 4800 = -820 kJmol-1 (1) |
|  | (b) | Bond enthalpies are mean values; specific bond enthalpies in reactants and products may differ (1)Water is a gas in above equation but liquid in standard enthalpy of combustion (1)[5] |
| **5.** | (a) | C4H10(g) + 6.5O2(g) 🡪 4CO2(g) + 5H2O(l) species and balanced (1) state symbols (1) |
|  | (b) | ΔH = ΣΔHf(p) - ΣΔHf(r) or correct cycle (1)= 4 x (-394) + 5 x (-286) – (-126) (1)= -2880 kJmol-1 (1)[5] |
| **6.** | (a) | 4C(s) + 2H2(g) + 2O2(g) 🡪 C4H4O4(s) species and balanced (1) state symbols (1) |
|  | (b) | ΔH = ΣΔHc(r) - ΣΔHc(p) or correct cycle (1)= 4 x (-393.5) + 2 x (-285.8) – (-1356) (1)= -789.6 kJmol-1 (1)[5] |
| **7.** | (a) | Ti(s) + 2Cl2(g) 🡪 TiCl4(l) species and balanced (1) state symbols (1) |
|  | (b) | ΔH = ΣΔHf(p) - ΣΔHf(r) or correct cycle (1)-232 = x + (-394) – (-912) so x = -232 + 394 – 912 (1)= -750 kJmol-1 (1)[5] |

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| **8.** | (a) | Ca(s) + F2(g) 🡪 CaF2(s) (1) |
|  | (b) | As below, but Ca should replace Mg and F should replace ClClear cycle showing species and states on each line (1)Correct states, multiples and shape (1) |
|  | (c) | ΔHf = 193 + 2(79) + 590 + 1150 + 2(-348) – 2611 (1) *(-2611 is the correct value for LE, not +193)*= -1216 kJmol-1 (1)*Note: using the incorrect value for LE (+193) given in the original assignment, answer = +1588 kJmol-1*[5] |
| **9.** | (a) | Lattice dissociation and hydration (1) balanced (1) |
|  | (b) | CaCl2(s) 🡪 Ca2+(aq) + 2Cl-(aq) (1) |
|  | (c) | 2(-364) + (-1650) – (-2237) (1)= -141 kJmol-1 (1) [5] |