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| **DEPARTMENT OF CHEMISTRY****FOURAH BAY COLLEGE****UNIVERSITY OF SIERRA LEONE**CHEM 311**PHYSICAL ANALYTICAL CHEMISTRY II****Unit 2 – Electroanalytical Techniques****CONTINUOUS ASSESSMENT****ASSIGNMENT**This assignment must be submitted no later than 2 pm on Friday March 23rd 2018You must submit this cover sheet with your assignment.Name: ……………………………………………………Admission No. ………………..Note: Unit 2 Continuous Assessment is worth 10% of the total marks for CHEM312Your score will be divided into three parts:Lecture and Tutorial Attendance 10%Assignment 40%Test 50% |

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| **1.** | (a) | (i) M1: potentiometric titration(ii) M2: amperometric titration(iii) M3: polarography(iv) M4: chronopotentiometry |
|  | (b) | (i) (ii)(iii)(iv) | M5: potential on y-axis, volume on x-axisM6: shallow-steep-shallow in either directionM7: current on y-axis, volume on x-axisM8: either V-shaped or obtuse angle flat before or after intersectionM9: current on y-axis, potential on x-axisM10: shallow-steep-shallow in either directionM11: potential on y-axis, time on x-axisM12: steep-shallow-steep in either direction  |
|  | (c) | 1. M13: constant current

M14: potential depends on concentration of electroactive species1. M15: constant potential

M16: current depends on concentration of electroactive species1. M17: diffusion current depends on concentration of electroactive species

M18: half-wave potential depends on identity of electroactive species1. M19: constant potential

M20: transition time depends on concentration of electroactive species |
|  | (d) | 1. M21: burette and beaker

M22: electrodes and voltmeter1. M23: burette and beaker

M24: electrodes and ammeter1. M25: dropping mercury electrode and non-polarisable electrode

M26: potentiometer and galvanometer1. M27: electrodes, digital timer and voltmeter

[max 25] |
| **2.** |  | M1: polarograph – apparatus used in polarographyM2: diagram showing DME and one other electrodeM3: diagram showing potentiometer, cell and galvanometerM4: polarogram – graph produced in polarographyM5: diffusion current shown on diagram as vertical distance between shallow sectionsM6: half-wave potential shown as mid-point of steep section[max 5] |
| **3.** | Describe the main features of a dropping mercury electrode and explain why it is useful in polarography.M1: mercury drops slowly through a capillary onto the anodeM2: the electrode is polarisable (it is difficult for current to actually flow)Any two from:M3: The dropping continually exposes fresh electrode to the electrolyteM4: This prevents the products of electrolysis from accumulatingM5: The droplets can be reproduced exactlyAny one from:M6: hydrogen has a high over-potential on mercury so the technique can be used in acid solutionsM7: mercury forms stable alloys (amalgams) with many metals and this can be used to analyse them [max 5] |

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| **4.** | (a) | M1: the current produced during electrolysisM2: when the process is limited by the rate at which the electroactive species is able to reach the electrode due to diffusion |
|  | (b) | M3: I = fMzMAFM4: so I = zMAFDM |
|  | (c) | M5: ID is reached when [M] at electrode is zero so I = zMAFDMM6: I = [M]zMAF |
|  | (d) | M7: m = mass per second = Vρt-1 so V = mtρ-1M8: V = M9: r3 = = M10: A = 4πr2 = 4πM11: A =  |
|  | (e) | M12: I = [M]zMAF so I = [M]zMFM13: let zMF be a constantM14: I [M] = [M]M15: average ID over lifetime of drop = total ID / τM16: ID [M] x  dt M17: dt = = M18: ID  |
|  | (f) | M19: based on linear diffusion but dropping electrode is spherically symmetricalM20: assumption only valid if thickness of diffusion layer is small compared to drop sizeM21: further adjustments need to be made which complicate the equationM22: relies on accurate values of DA being available, these are only available at infinite dilutionM23: proportionality different for different polarographs due to variations in m and τM24 and M25 available for quality of language [25] |
| **5.** | (a) | M1: E = Eo +  |
|  | (b) | M2: I = k√Do([O] – [O]e) so [O]e = M3: diffusion current reached when [O]e = 0 so ID = k√Do[O]M4: so [O]e = M5: I = k√Dr([R]e – [R]) but [R] = 0 so I = k√Dr[R]e so [R]e =  |
|  | (c) | M6: = = M7: E = Eo + = Eo + = Eo + M8: At V1/2, I = ID/2 so = = 1 so Eo +  |
|  | (d) | M9: it is valid provided that Dr and Do are similar, which is usually the caseM10: but the Eo value is specific to a mercury electrode[10] |
| **6.** | (a) | M1: Mx + pLy 🡪 [MLp]x+pyM2: Kstab =  |
|  | (b) | (i) | M3: E1/2 = half-wave potential and Eo = standard electrode potential and R = molar gas constant and T = absolute temperature and F = Faraday constant and n = number of electrons gained or lost during the half-reaction (any 5) |
|  |  | (ii) | M4: measure E1/2 at several different concentrations of ligand and plot E1/2 against ln[Ly]M5: gradient of line = and y-intercept = Eo - lnKstab[5] |
| **7.** | (a) | M1: M3:M5:  | M2: Reducible species in beaker, concentration decreases at it reacts, no change in activity after equivalence point, example Pb2+ against SO42-M4: Reducible species in burette, concentration low until equivalence point, then increases, no change in activity after equivalence point, example SO42- against Pb2+M6: Reducible species in burette and beaker, concentration decreases until equivalence point due to decrease in concentration of titrand, then increases due to increase in concentration of titrant, example titration of Pb2+ against Cr2O72-Any 2 of M2, M4, M5 [5] |
| **8.** | (a) | M1: amperometric measures current vs volume at constant potentialM2: potentiometric measures potential vs volume at constant current |
|  | (b) | In potentiometric titrations, a graph of ΔE/ΔV against V is more useful than a graph of E against V. Explain why, and sketch the shape of a graph of ΔE/ΔV against V.M3: point of maximum gradient is easier to see on an ΔE/ΔV vs Vol graphM4: clear spike either up or downM5: y-axis dE/dV and x-axis Vol[5] |

**TOTAL 85 MARKS**