

## UNIT 5B PRACTICE QUIZ 2 – DISPLACEMENT REACTIONS AND THE REACTIVITY SERIES

Do not answer these questions on this document.

Write your answers on a sheet of paper; then click on the answer sheet provided at the end of the questions.

Use your Periodic Table and the reactivity series in your course notes.

You will also need the following information:

<b>1.</b>	Which of the following metals will react the most rapidly with water?	
	A	Barium
	B	Beryllium
	C	Calcium
	D	Magnesium
	E	Strontium
1		

<b>2.</b>	When metals react with water,	
	A	The metal is oxidised and H is reduced
	B	The metal is oxidised and O is reduced
	C	The metal is reduced and H is oxidised
	D	The metal is reduced and O is oxidised
1		

<b>3.</b>	How would you expect lead to react with acids?	
	A	Explosively
	B	Rapidly
	C	Slowly (reaction happens but it barely visible)
	D	Not at all
1		

<b>4.</b>	How would you expect silver to react with acids?	
	A	Explosively
	B	Rapidly
	C	Slowly (reaction happens but it barely visible)
	D	Not at all
1		

UNIT 5B – CHEMICAL REACTIONS II – REDOX REACTIONS

<b>5.</b>	If some iron metal is added to a solution of copper chloride,	
	A	nothing will happen
	B	The iron will be oxidized and the copper will be reduced
	C	The copper will be oxidised and the iron will be reduced
	D	The iron will be oxidized and the chlorine will be reduced
<b>1</b>		

<b>6.</b>	If some iron metal is added to a solution of zinc oxide,	
	A	nothing will happen
	B	The iron will be oxidised and the zinc will be reduced
	C	The zinc will be oxidised and the iron will be reduced
	D	The iron will be oxidized and the oxygen will be reduced
<b>1</b>		

<b>7.</b>	The element zinc could be extracted from its main ore by	
	A	carbon but not hydrogen
	B	hydrogen but not carbon
	C	either hydrogen or carbon
	D	neither hydrogen nor carbon
<b>1</b>		

<b>8.</b>	The element silver could be reacted from its main ore by	
	A	either carbon or hydrogen
	B	hydrogen but not carbon
	C	carbon but not hydrogen
	D	neither hydrogen nor carbon
<b>1</b>		

<b>9.</b>	Copper reacts with mercury oxide as follows: $\text{Cu} + \text{HgO} \rightarrow \text{CuO} + \text{Hg}$ In this reaction	
	A	Hg is oxidized and O is reduced
	B	Cu is oxidized and O is reduced
	C	Hg is oxidized and Cu is reduced
	D	Cu is oxidised and Hg is reduced
<b>1</b>		

UNIT 5B – CHEMICAL REACTIONS II – REDOX REACTIONS

<b>10.</b>	The reaction in question 9 shows that	
	A	mercury is more reactive than copper
	B	copper is more reactive than mercury
	C	mercury is more reactive than oxygen
	D	oxygen is more reactive than mercury
<b>1</b>		

[Go to the answer sheet](#)