



General Certificate of Education

Chemistry 1421

CHEM1 Foundation Chemistry

Mark Scheme

2010 examination - January series

Mark schemes are prepared by the Principal Examiner and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation meeting attended by all examiners and is the scheme which was used by them in this examination. The standardisation meeting ensures that the mark scheme covers the candidates' responses to questions and that every examiner understands and applies it in the same correct way. As preparation for the standardisation meeting each examiner analyses a number of candidates' scripts: alternative answers not already covered by the mark scheme are discussed at the meeting and legislated for. If, after this meeting, examiners encounter unusual answers which have not been discussed at the meeting they are required to refer these to the Principal Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of candidates' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this Mark Scheme are available to download from the AQA Website: www.aqa.org.uk

Copyright © 2010 AQA and its licensors. All rights reserved.

COPYRIGHT

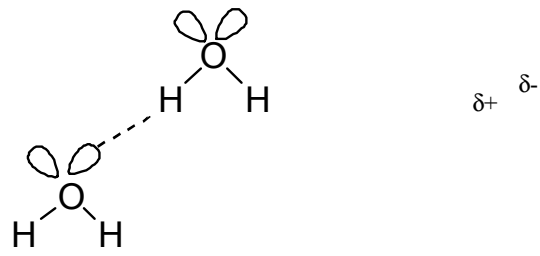
AQA retains the copyright on all its publications. However, registered centres for AQA are permitted to copy material from this booklet for their own internal use, with the following important exception: AQA cannot give permission to centres to photocopy any material that is acknowledged to a third party even for internal use within the centre.

Set and published by the Assessment and Qualifications Alliance.

1	(d)		Lower Two/pair of electrons in (3)p orbital or implied repel (each other)	1 1 1	If not lower CE = 0/3 If blank mark on Allow does not increase Not 2p M3 dependent upon a reasonable attempt at M2
1	(e)		Boron/B or oxygen/O/ O ₂	1	

Question	Part	Sub Part	Marking Guidance	Mark	Comments
2	(a)	(i)	$M_r = 132.1$	1	132
			0.0238	1	Allow 0.024 Allow 0.0237 Penalise less than 2 sig fig once in (a)
2	(a)	(ii)	0.0476	1	0.0474-0.0476 Allow (a) (i) x 2
2	(a)	(iii)	1.21	1	Allow consequential from (a) (ii) ie allow (a) (ii) x 1000 / 39.30 Ignore units even if wrong
2	(b)		$\frac{34 \times 100}{212.1}$	1	Allow mass or Mr of desired product times one hundred divided by total mass or Mr of reactants/products If 34/212.1 seen correctly award M1
			= 16.0(3)%	1	Allow 16% 16 scores 2 marks
2	(c)		100(%)	1	Ignore all working
2	(d)		$PV = nRT$ or $n = \frac{PV}{RT}$	1	If rearranged incorrectly lose M1 and M3
			$n = \frac{100000 \times 1.53 \times 10^{-2}}{8.31 \times 310}$	1	M2 for mark for converting P and T into correct units in any expression
			= 0.59(4)	1	Allow 0.593 M3 consequential on transcription error only not on incorrect P and T

2	(e)		(Na ₂ SO ₄) (44.1%)	H ₂ O 55.9%	1	M1 is for 55.9
			44.1/142.1 0.310 =1	55.9/18 3.11 =10	1	Alternative method gives 180 for water part = 2 marks
			x = 10		1	X = 10 = 3 marks 10.02 = 2 marks

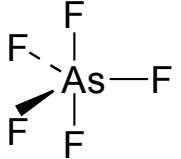
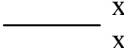
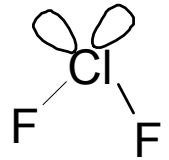
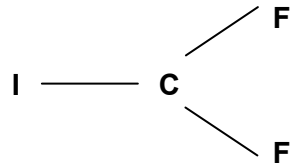
Question	Part	Sub Part	Marking Guidance	Mark	Comments
3	(a)		Hydrogen/H bonds	1	Not just hydrogen
			van der Waals/vdw/ dipole-dipole/London/temporarily induced dipole/dispersion forces	1	Not just dipole
3	(b)			3	M1 for partial charges as indicated in diagram (correct minimum) M2 for all four lone pairs M3 for H bond from the lp to the H (δ^+) on the other molecule Lone pair on hydrogen CE = 0 OHO CE = 0 If only one molecule of water shown CE = 0
3	(c)		Hydrogen bonds/IMF (in water) stronger OR IMF / VDW / dipole-dipole forces (in H_2S) are weaker OR H bonding is the strongest IMF	1	Ignore energy references Comparison must be stated or implied
3	(d)		Atoms/molecules get larger/more shells/more electrons/ more surface area	1	Not heavier/greater Mr
			therefore increased <u>Van der Waals/IMF</u> forces	1	Ignore references to dipole-dipole forces

3	(e)		Dative (covalent)/ coordinate	1	If not dative/coordinate CE = 0/2 If covalent or blank read on
			(Lone) pair/both electrons/two electrons on O(H ₂) donated (to H ⁺) OR pair/both electrons come from O(H ₂)	1	Explanation of a coordinate bond specific to oxygen or water required Not just H ⁺ attracted to lone pair since that is nearer to a H bond
3	(f)		ionic	1	if not ionic CE = 0
			oppositely charged <u>ions</u> /+ and – <u>ions or particles</u>	1	atoms or molecules loses M2 and M3
			ions attract <u>strongly</u> OR strong/many (ionic) bonds must be broken	1	S ⁻ loses M2 Reference to IMF loses M2 and M3

Question	Part	Sub Part	Marking Guidance	Mark	Comments
4	(a)	(i)	single (C-C) bonds <u>only</u> / no double (C=C) bonds	1	Allow all carbon atoms bonded to four other atoms Single C-H bonds only = 0 C=H CE
			C and H (atoms) <u>only/purely/solely/entirely</u>	1	Not consists or comprises Not completely filled with hydrogen CH molecules = CE Element containing C and H = CE
4	(a)	(ii)	C_nH_{2n+2}	1	<i>Formula only</i> C_xH_{2x+2}
4	(b)	(i)	$C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$	1	Accept multiples Ignore state symbols
4	(b)	(ii)	gases produced are greenhouse gases/contribute to Global warming/effect of global warming/climate change	1	Allow CO ₂ or water is greenhouse gas/causes global warming Acid rain/ozone CE = 0
4	(c)		carbon	1	Allow C Allow soot
4	(d)	(i)	$C_9H_{20} \rightarrow C_5H_{12} + C_4H_8$	1	Accept multiples
			OR $C_9H_{20} \rightarrow C_5H_{12} + 2C_2H_4$		
4	(d)	(ii)	Plastics, polymers	1	Accept any polyalkene / haloalkanes / alcohols
4	(d)	(iii)	so the <u>bonds</u> break OR because the <u>bonds</u> are strong	1	IMF mentioned = 0
4	(e)	(i)	1,4-dibromo-1-chloropentane / 1-chloro-1,4-dibromopentane	1	Ignore punctuation
4	(e)	(ii)	Chain/position/positional	1	Not structural or branched alone

Question	Part	Sub Part	Marking Guidance	Mark	Comments
5	(a)		<p><u>Average/mean mass of (1) atom(s) (of an element)</u> 1/12 mass of one atom of ^{12}C</p> <p>OR</p> <p><u>(Average) mass of one mole of atoms</u> 1/12 mass of one mole of ^{12}C</p> <p>OR</p> <p><u>(Weighted) average mass of all the isotopes</u> 1/12 mass of one atom of ^{12}C</p> <p>OR</p> <p>Average mass of an atom/isotope compared to C-12 on a scale in which an atom of C-12 has a mass of 12</p>	1 1	<p>If moles and atoms mixes Max = 1</p> <p>This expression = 2 marks</p>
5	(b)		<p>d block</p> <p>$[\text{Ar}] 3\text{d}^2 4\text{s}^2$</p> <p>27</p>	1 1 1	<p>Allow 3d/D</p> <p>Other numbers lose M1</p> <p>Ignore transition metals</p> <p>Can be written in full</p> <p>Allow subscripts</p> <p>3d^2 and 4s^2 can be in either order</p>

5	(c)		$\frac{(90 \times 9) + (91 \times 2) + (92 \times 3) + (94 \times 3)}{17}$ (= 1550) (or Σ their abundances)	1	If one graph reading error lose M1 and allow consequential M2 and M3. If 2 GR errors penalise M1 and M2 but allow consequential M3 If not 17 or Σ their abundances lose M2 and M3 91.2 = 3 marks provided working shown. M4 -allow nearest consequential element from M3 accept Zr in any circumstance
			=91.2	1	
			Zr/ Zirconium	1	
5	(d)		High energy electrons/bombarded or hit with electrons	1	accept electron gun If not 90 lose M3 and M4 If charge is wrong on 90 isotope lose M3 only Accept any symbol in place of Z Allow lightest
			knocks out electron(s) (to form ions)	1	
			$Z^+ = \underline{90}$ deflected most	1	
			since lowest mass/lowest m/z	1	
5	(e)		(ions hit detector and) cause current/(ions) accept electrons/cause electron flow	1	QWC
			bigger current = more of that isotope/current proportional to abundance	1	Implication that current depends on the number of ions

Question	Part	Sub Part	Marking Guidance	Mark	Comments
6				1	<p>Mark M1 – M5 independently M1 for 5 bond pairs around As Do not penalise A for As or F1 for F</p> <p>Allow trigonal dipyramid</p> <p>M3 for 2 bond pairs to F and 2 lone pairs Lone pairs can be shown as lobes with or without electrons or as xx or</p> <p style="text-align: center;">  </p> <p>Bent-linear = contradiction Do not allow trigonal</p> <p>Not just triangular</p>
			trigonal / triangular bipyramid(al)	1	
				1	
			Bent / V shape / non-linear / triangular / angular	1	
			104° - 106° (For candidates who thought this was ClF ₂ ⁺ which contained iodine allow	1	
					
Trigonal / triangular <u>planar</u>					
120°					