



General Certificate of Education

Chemistry 1421

CHEM2 Chemistry in Action

**Post-Standardisation Mark
Scheme**

2010 examination - January series

Mark schemes are prepared by the Principal Examiner and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation meeting attended by all examiners and is the scheme which was used by them in this examination. The standardisation meeting ensures that the mark scheme covers the candidates' responses to questions and that every examiner understands and applies it in the same correct way. As preparation for the standardisation meeting each examiner analyses a number of candidates' scripts: alternative answers not already covered by the mark scheme are discussed at the meeting and legislated for. If, after this meeting, examiners encounter unusual answers which have not been discussed at the meeting they are required to refer these to the Principal Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of candidates' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

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Q	Part	Sub Part	Marking Guidance	Mark	Comments
1	a	i	Reducing agent OR Reduce(s) (WO ₃ / tungsten oxide) OR electron donor OR to remove <u>oxygen</u> (from WO ₃ / tungsten oxide or to form water);	1	
1	a	ii	WO ₃ + 3H ₂ → W + 3H ₂ O	1	Or multiples
1	a	iii	<i>One from</i> H ₂ is <ul style="list-style-type: none"> • explosive • flammable or inflammable • easily ignited 	1	Ignore reference to pressure or temperature
1	b	i	Addition OR (catalytic) hydrogenation OR Reduction	1	Ignore “electrophilic” Penalise “nucleophilic addition”
1	b	ii	Geometric(al) OR cis/trans OR E Z OR E/Z	1	
1	c	i	(If any factor is changed which affects an equilibrium), the position of <u>equilibrium</u> will <u>shift / move / change/ respond / act</u> so as <u>to oppose the change</u> . OR (When a system/reaction in equilibrium is disturbed), the <u>equilibrium shifts / moves</u> in a direction which tends <u>to reduce the disturbance</u>	1	A variety of wording will be seen here and the key part is the last phrase and must refer to <u>movement of the equilibrium</u> . QoL

1	c	ii	<p>M1 – Statement of number of moles / molecules There are <u>more moles / molecules</u> (of gas) on the left / of reactants OR <u>fewer moles / molecules</u> (of gas) on the right./ products OR there are <u>4 moles / molecules</u> (of gas) on the left <u>and 2 moles / molecules</u> on the right.</p> <p>M2 – Explanation of response / movement in terms of pressure <u>Increase in pressure is opposed</u> (or words to that effect) OR <u>pressure is lowered</u> by a shift in the equilibrium (from left) <u>to right</u> / favours forward reaction.</p>	2	<p>Ignore "volumes" for M1</p> <p>Mark independently</p>
1	d		<p>$\sum B(\text{reactants}) - \sum B(\text{products}) = \Delta H$ (M1) OR <u>Sum of bonds broken</u> – <u>Sum of bonds formed</u> = ΔH (M1)</p> <p>$B(\text{H-H}) + \frac{1}{2}B(\text{O=O}) - 2B(\text{O-H}) = -242$ (M1)</p> <p>$B(\text{H-H}) = -242 - \frac{1}{2}(+496) + 2(+463)$ (this scores M1 and M2)</p> <p>$B(\text{H-H}) = (+)436$ (kJ mol^{-1}) (M3)</p> <p>Award 1 mark for – 436</p> <p>Candidates may use a cycle and gain full marks.</p>	3	<p>M1 could stand alone</p> <p><u>Award full marks for correct answer.</u></p> <p>Ignore units.</p> <p>Two marks can score with an arithmetic error in the working.</p>

Q	Part	Sub Part	Marking Guidance	Mark	Comments
2	a		<u>Heat (energy) change at constant pressure</u>	1	Ignore references to standard conditions, but credit specified pressure.
2	b		The <u>enthalpy change</u> / <u>heat (energy) change</u> (at constant pressure) in a reaction is independent of the route / path taken (and depends only on the initial and final states)	1	
2	c		$\Delta H + 963 = -75 - 432$ OR $\Delta H + 963 = -507$ (M1) $\Delta H = -75 - 432 - 963$ (M1 and M2) $\Delta H = \underline{-1470}$ (kJ mol ⁻¹) Award 1 mark for + 1470	3	<u>Award full marks for correct answer</u> Ignore units. Ignore numbers on the cycle M1 and M2 can score for an arithmetic error

Q	Part	Sub Part	Marking Guidance	Mark	Comments
3	a		NaBr ONLY	1	Penalise incorrect case or additional formulae. Ignore names
3	b		NaF ONLY	1	Penalise incorrect case or additional formulae. Ignore names
3	c		<u>ONLY one</u> from either NaF OR NaCl	1	Penalise incorrect case or additional formulae. Ignore names
3	d		NaI ONLY	1	Penalise incorrect case or additional formulae. Ignore names

Q	Part	Sub Part	Marking Guidance	Mark	Comments
4	a		Antacid OR to neutralise acidity OR eases indigestion	1	Credit suitable reference to indigestion or to laxative or to relief of constipation
4	b		M1 Decrease in T decreases the <u>energy</u> of the <u>particles / ions / H⁺ / molecules</u> M2 (also scores M1) <u>Decrease in the number of / less particles / ions / H⁺ / molecules with $E \geq E_{\text{Act}}$ or $E \geq$ minimum energy to react</u> M3 <u>Few(er) / Less effective / productive / successful collisions</u>	3	In M1 and M2 , credit “atoms” but ignore “calcium carbonate”, ignore “calcium”, ignore any ion formula except H ⁺ QoL
4	c	i	Strontium has a higher melting point than barium, because Correct reference to size of cations/proximity of electrons M1 (For Sr) delocalised <u>electrons closer to cations / positive ions / atoms / nucleus</u> OR <u>cations / positive ions / atoms are smaller</u> OR <u>cation / positive ion / atom or it has fewer (electron) shells / levels</u> Relative strength of metallic bonding M2 (Sr) has <u>stronger</u> attraction between the <u>cations / positive ions / atoms / nucleus</u> and the delocalised <u>electrons</u> OR <u>stronger metallic bonding</u> (assume argument refers to Sr but accept converse argument for Ba)	2	Ignore general Group 2 statements Penalise M1 if Sr or Ba is said to have <u>more or less</u> delocalised electrons Ignore reference to shielding CE = 0 for reference to molecules or intermolecular forces or covalent bonds Ignore “Van der Waals forces (between atoms)” but penalise if “between molecules”
4	c	ii	$\text{Sr} + 2\text{H}_2\text{O} \longrightarrow \text{Sr}(\text{OH})_2 + \text{H}_2$	1	Or multiples
4	d	i	$2\text{Mg} + \text{TiCl}_4 \longrightarrow 2\text{MgCl}_2 + \text{Ti}$	1	Or multiples

4	d	ii	It or MgSO_4 is <u>soluble</u> OR forms <u>a solution</u> (and is washed away) OR <u>dissolves</u>	1	Credit reference to MgSO_4 being the most soluble Group 2 sulfate. Ignore "disappears"
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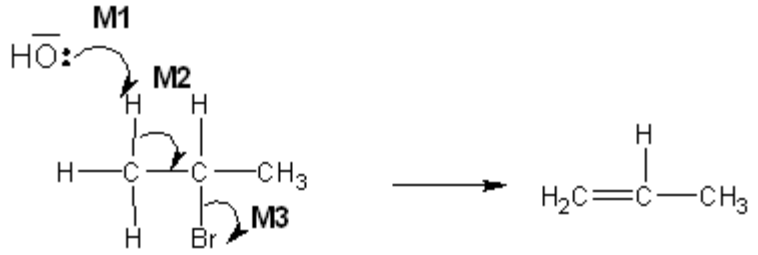
Q	Part	Sub Part	Marking Guidance	Mark	Comments
5	a	i	Oxidation OR Oxidised ONLY	1	
5	a	ii	Any one from <ul style="list-style-type: none"> to provide / overcome activation energy to provide the minimum energy to make the reaction go / start 	1	NOT simply to increase the (initial) reaction rate.
5	a	iii	The reaction is exothermic OR releases heat (energy)	1	
5	a	iv	M1 Catalysts provide an alternative route / pathway OR an alternative mechanism OR (in this case) surface adsorption occurs (or a description of adsorption) M2 Lowers the activation energy OR of lower activation energy	2	Ignore reference to "surface" alone
5	b		M1 The (forward) reaction is exothermic OR the (forward) reaction releases heat OR The reverse reaction is endothermic or absorbs heat M2 – Direction of change N.B. M2 depends on correct M1 At lower temperatures, <ul style="list-style-type: none"> the equilibrium yield of NO₂ is greater more NO₂ is formed equilibrium shifts (left) to right (equilibrium) favours the forward reaction (OR converse for higher temperatures)	2	

5	c		NO ₂ (+) 4		3	
			NO ₃ ⁻ (+) 5			
			HNO ₂ (+) 3			

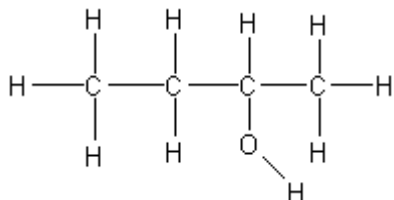
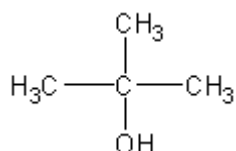
Q	Part	Sub Part	Marking Guidance	Mark	Comments	
6	a		<u>Functional group</u> (isomerism)	1		
6	b		<p>M1 Tollens' (reagent) (<i>Credit ammoniacal silver nitrate OR a description of making Tollens'</i>) (<i>Ignore either AgNO₃ or [Ag(NH₃)₂⁺]</i> or "the silver mirror test" on their own, but mark M2 and M3) M2 <u>silver mirror</u> OR <u>black solid/precipitate</u> (<i>NOT silver precipitate</i>)</p> <p>M3 (stays) colourless or no change or no reaction</p>	<p>M1 Fehling's (solution) or Benedict's solution (<i>Ignore Cu²⁺(aq) or CuSO₄ on their own, but mark on to M2 and M3</i>)</p> <p>M2 <u>Red solid/precipitate</u> (<i>Credit orange or brown solid</i>)</p> <p>M3 (stays) blue or no change or no reaction</p>	3	<p>No reagent, CE=0</p> <p>Allow the following alternatives M1 (acidified) potassium dichromate(VI) (solution) M2 (turns) green M3 (stays) orange / no change OR M1 (acidified) potassium manganate(VII) (solution) M2 (turns) colourless M3 (stays) purple / no change</p> <p>For M3 Ignore "nothing (happens)" Ignore "no observation"</p>
6	c		(Both have) C=O OR a carbonyl (group)	1		
6	d	i	(Free-) <u>radical substitution</u> ONLY	1	Penalise "(free) radical mechanism"	

6	d	ii	<p>Initiation $\text{Cl}_2 \longrightarrow 2\text{Cl}\cdot$</p> <p>First propagation $\text{Cl}\cdot + \text{CH}_3\text{CH}_2\text{CH}_3 \longrightarrow \cdot\text{CH}_2\text{CH}_2\text{CH}_3 + \text{HCl}$ OR C_3H_8</p> <p>Second propagation $\text{Cl}_2 + \cdot\text{CH}_2\text{CH}_2\text{CH}_3 \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{Cl} + \text{Cl}\cdot$ OR $\text{C}_3\text{H}_7\text{Cl}$</p> <p>Termination (must make C_6H_{14}) $2 \cdot\text{CH}_2\text{CH}_2\text{CH}_3 \longrightarrow \text{C}_6\text{H}_{14}$ or $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$</p>	4	<p>Penalise absence of dot once only.</p> <p>Penalise incorrect position of dot on propyl radical once only.</p> <p>Penalise $\text{C}_3\text{H}_7\cdot$ once only</p> <p>Accept $\text{CH}_3\text{CH}_2\text{CH}_2\cdot$ with the radical dot above / below / to the side of <u>the last carbon</u>.</p> <p>Use of the secondary free radical might gain 3 of the four marks</p>
6	e		<p>$M_r = \underline{44.06352}$ (for propane) $M_r = \underline{43.98982}$ (for carbon dioxide)</p> <p>M1 a correct value for <u>both</u> of these <u>M_r values</u>.</p> <p>M2 a statement or idea that <u>two peaks</u> appear (in the mass spectrum) OR <u>two molecular ions</u> are seen (in the mass spectrum).</p>	2	Mark independently

Q	Part	Sub Part	Marking Guidance	Mark	Comments
7	a	i	<p><u>Nucleophilic substitution</u></p> <p>M1 must show an arrow from the lone pair of electrons on the oxygen atom of the negatively charged hydroxide ion to the central C atom.</p> <p>M2 must show the movement of a pair of electrons from the C-Br bond to the Br atom. Mark M2 independently.</p> <p>Award full marks for an S_N1 mechanism in which M1 is the attack of the hydroxide ion on the intermediate carbocation.</p>	1 2	<p>Penalise M1 if covalent KOH is used</p> <p>Penalise M2 for formal charge on C or incorrect partial charges</p> <p>Penalise once only for a line and two dots to show a bond.</p> <p>Max 1 mark for the mechanism for the wrong reactant and/or "sticks"</p> <p>Ignore product</p>
7	a	ii	2-bromopropane ONLY	1	
7	a	iii	<p><u>Polar C–Br</u> OR <u>polar carbon–bromine bond</u> OR <u>dipole on C–Br</u></p> <p>OR C atom of <u>carbon–bromine bond</u> is δ+ / electron deficient OR C–Br δ+ (δ–)</p> <p>(Credit <u>carbon–halogen bond</u> as an alternative to <u>carbon–bromine bond</u>.)</p>	1	<p>It must be clear that the discussion is about the carbon atom of the C–Br bond. NOT just reference to a polar molecule.</p> <p>Ignore X for halogen</p>

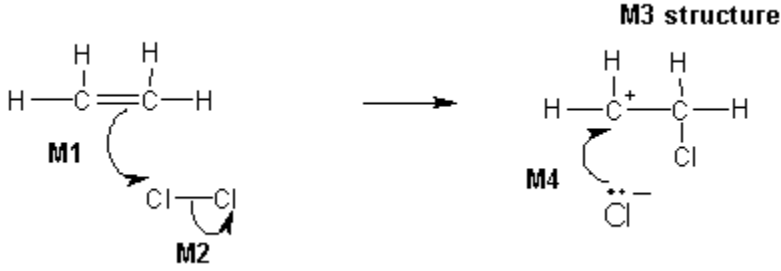
7	b		<p><u>Elimination</u></p>  <p>M1 must show an arrow from the lone pair on oxygen of a negatively charged hydroxide ion to the correct H atom M2 must show an arrow from the correct C-H bond to the C-C bond and should only be awarded if an attempt has been made at M1 M3 is independent.</p> <p>Award full marks for an E1 mechanism in which M2 is on the correct carbocation.</p>	1 3	<p>Credit “base elimination” but NOT “nucleophilic elimination” No other prefix.</p> <p><u>Mechanism</u> Penalise M1 if covalent KOH</p> <p>Penalise M3 for formal charge on C or incorrect partial charges</p> <p>Penalise once only for a line and two dots to show a bond.</p> <p>Max 2 marks for the mechanism for wrong reactant and/or “sticks”</p> <p>Ignore product</p>
7	c		<p><i>Any one condition from this list to favour elimination;</i></p> <ul style="list-style-type: none"> • <u>alcohol(ic) / ethanol(ic)</u> (solvent) • <u>high concentration</u> of KOH / alkali / hydroxide OR <u>concentrated</u> KOH / hydroxide • high temperature or hot or heat under reflux or T = 78 to 100°C 	1	<p>Apply the list principle</p> <p>Ignore “aqueous”</p> <p>Ignore “excess”</p>
7	d	i	<u>Addition</u> (polymerisation) ONLY	1	Penalise “additional”
7	d	ii	<u>But-2-ene</u> ONLY (hyphens not essential)	1	<p>Ignore references to cis and trans or E/Z</p> <p>Ignore butene</p>

Q	Part	Sub Part	Marking Guidance	Mark	Comments
8	a	i	$2\text{CuFeS}_2 + 2\text{SiO}_2 + 4\text{O}_2 \longrightarrow \text{Cu}_2\text{S} + 2\text{FeSiO}_3 + 3\text{SO}_2$	1	
8	a	ii	Acid rain OR an effect either from acid rain or from an acidic gas in the atmosphere	1	
8	a	iii	SO_2 could be used to make H_2SO_4 OR to make gypsum / plaster or $\text{CaSO}_4(\text{xH}_2\text{O})$	1	
8	b		$\text{Cu}_2\text{S} + 2\text{O}_2 \longrightarrow 2\text{CuO} + \text{SO}_2$	1	Or multiples Ignore state symbols
8	c		$2\text{CuO} + \text{C} \longrightarrow 2\text{Cu} + \text{CO}_2$ OR $\text{CuO} + \text{C} \longrightarrow \text{Cu} + \text{CO}$	1	Or multiples Ignore state symbols
8	d	i	<i>Any one from the following two ONLY</i> <ul style="list-style-type: none"> • <u>(Scrap) iron is cheap</u> • <u>Low energy requirement</u> 	1	Apply the list principle Not "less energy"
8	d	ii	$\text{Fe} + \text{Cu}^{2+} \longrightarrow \text{Fe}^{2+} + \text{Cu}$	1	Or multiples Ignore state symbols

Q	Part	Sub Part	Marking Guidance	Mark	Comments
9	a		<p>M1 <u>Displayed formula</u> for butan-2-ol</p>  <p>M2 Alcohol X is</p>  <p>M3 Alcohol Y is named <u>(2)-methylpropan-1-ol</u> ONLY</p>	3	<p>M1 displayed formula <u>must</u> have all bonds drawn out, including the O—H but ignore angles</p> <p>Penalise “sticks”</p> <p>M2 structure must be clearly identifiable as 2-methylpropan-2-ol and may be drawn in a variety of ways.</p> <p>M3 <u>must be correct name</u>, but ignore structures</p>
9	b		<p>M1 The infrared spectrum shows an <u>absorption / peak in the range 3230 to 3550 (cm⁻¹)</u>(which supports the idea that an alcohol is present)</p> <p>M2 Reference to the ‘fingerprint region’ or below 1500 (cm⁻¹)</p> <p>M3 <u>Match with</u> or <u>same as</u> known sample / database spectra</p> <p>OR alternatively</p> <p>M2 Run infrared spectra (of the alcohols)</p> <p>M3 Find which one <u>matches</u> or is the <u>same as</u> this spectrum.</p>	3	<p>In M1, allow the words “dip”, “spike”, “low transmittance” and “trough” as alternatives for absorption.</p> <p>Check the spectrum to see if alcohol OH is labelled and credit.</p>

9	c	<p>M1 balanced equation $\text{C}_6\text{H}_{12}\text{O}_6 \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH} + 2\text{CO}_2 + \text{H}_2\text{O}$ or $\text{C}_4\text{H}_9\text{OH}$</p> <p>M2 Any one from</p> <ul style="list-style-type: none"> • <u>excess/adequate/sufficient/ correct amount of /enough/plenty / a good supply</u> of oxygen or air • good mixing of the fuel and air/oxygen <p>M3 $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH} + 6\text{O}_2 \longrightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$ or $\text{C}_4\text{H}_9\text{OH}$</p> <p>M4 A biofuel is a fuel produced <u>from</u> (renewable) <u>biological</u> (re)source(s) OR (renewable)_(re)source(s) <u>from</u> (a specified) <u>plant(s) /fruit(s) /tree(s)</u></p>	4	<p>Or multiples for M1 and M3</p> <p>In M1 and M3 penalise use of $\text{C}_4\text{H}_{10}\text{O}$ or butan-2-ol once only</p> <p>For M2, do <u>not</u> accept simply “oxygen” or “air” alone Ignore reference to “temperature”</p> <p>In M4 Ignore references to “carbon neutral” Ignore “sugar” and “glucose”</p>
9	d	<p>M1 butan-1-ol is a <u>primary or 1°</u> (alcohol)</p> <p>M2 <u>Displayed formula</u> (ONLY) for butanal $\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$</p> <p>M3 <u>Displayed formula</u> (ONLY) for butanoic acid $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$</p> <p>M4 Oxidation (oxidised) OR Redox</p> <p>M5 <u>orange to green</u></p>	5	<p>M2 and M3 displayed formula must have all bonds drawn out including the O—H but ignore angles.</p> <p>If butanal and butanoic acid formulae are <u>both</u> correctly given but not displayed, credit one mark out of two.</p> <p>Both colours required for M5 Ignore states</p>

Q	Part	Sub Part	Marking Guidance	Mark	Comments
10	a		<p>M1 Cl₂ (provides the pale green colour)</p> <p>M2 NaOH reacts <u>with the acid(s) / the HCl / the HClO / H⁺</u></p> <p>M3 <u>requires a correct answer in M2</u></p> <p>Equilibrium shifts (from left) <u>to right</u> OR wtte</p>	3	<p>M1 <u>requires the formula</u></p> <p>Ignore “reacts with the products”</p> <p>Ignore “reacts with chloride ion”</p> <p>Ignore “reacts with chlorine”</p>
10	b		<p>M1 A reducing agent is an <u>electron donor</u> OR (readily) <u>loses / gives away electrons</u></p> <p>M2 Cl₂ + 2e⁻ → 2Cl⁻</p> <p>For M3 and M4, iodide ions are stronger reducing agents than chloride ions, because</p> <p>M3 Relative size of ions / atomic radius / ionic radius <u>iodide ions</u> are <u>larger</u> / have more (electron) shells / levels than chloride ions (or converse for chloride ion) OR <u>electron(s) to be lost/outer shell/level is further</u> from the nucleus (or converse for chloride ion) OR greater / more shielding</p> <p>M4 Strength of attraction for <u>electron(s) being lost</u> <u>Electron(s) lost</u> from an iodide ion is <u>less strongly held by the nucleus</u> compared with that lost from a chloride ion</p> <p>(assume argument refers to iodide ions but accept converse argument for chloride ions)</p>	4	<p>Penalise M1 if “electron pair donor”</p> <p>Ignore state symbols in M2 Accept no charge on the electron</p> <p>Credit the electrons being lost on the RHS</p> <p>M3 and M4 must be comparative and should refer to electrons.</p> <p>For M3 insist on “<u>iodide ions</u>”</p>

10	c	<p>M1 $2\text{Cl}_2 + 2\text{H}_2\text{O} \longrightarrow 4\text{HCl} + \text{O}_2$</p> <p>M2 <u>silver chloride</u> ONLY</p> <p>M3 The solid / precipitate would dissolve OR is soluble OR (It) forms a (colourless) solution</p>	3	<p>Or multiples</p> <p><u>M2 requires a name</u></p> <p>Mark M3 independently</p> <p>Ignore "disappears"</p>
10	d	<p><u>Electrophilic addition</u></p> <p><i>Mechanism:</i></p> <div style="text-align: center;">  </div> <p>M1 must show an arrow from the double bond towards one of the Cl atoms on a Cl-Cl molecule.</p> <p>M2 must show the breaking of the Cl-Cl bond.</p> <p>M3 is for the structure of the carbocation with Cl substituent.</p> <p>M4 must show an arrow from the lone pair of electrons on a negatively charged chloride ion towards the positively charged carbon atom.</p>	1 4	<p>M2 Penalise partial charges if wrong way around, otherwise ignore</p> <p>Max 3 marks for the mechanism for wrong reactant and/or "sticks" (wrong reactant could be HBr or Br₂ or incorrect alkene)</p>

General principles applied to marking CHEM2 papers by CMI+ for January 2010

It is important to note that the guidance given here is generic and specific variations may be made at individual standardising meetings in the context of particular questions and papers.

A. The “List principle” and the use of “ignore” in the mark scheme

If a question requires **one** answer and a candidate gives two answers, no mark is scored if one answer is correct and one answer is incorrect. There is no penalty if both answers are correct.

N.B. Certain answers are designated in the mark scheme as those which the examiner should “Ignore”. These answers are not counted as part of the list and should be ignored and will not be penalised.

B. Incorrect case for element symbol

The use of an incorrect case for the symbol of an element should be penalised **once only** within a clip. For example, penalise the use of “h” for hydrogen, “CL” for chlorine or “br” for bromine.

C. Spelling

In general

- The names of chemical compounds and functional groups **must be spelled correctly** to gain credit.
- Phonetic spelling may be acceptable for some chemical terminology.

N.B. Some terms may be required to be spelled correctly as part of the “Quality of Language” (QoL) marking.

D. Equations

In general

- Equations **must** be balanced.
- When an equation is worth two marks, one of the marks in the mark scheme will be allocated to one or more of the reactants or products. This is independent of the equation balancing.
- State symbols are generally ignored, unless specifically required in the mark scheme.

E. Reagents

The guiding principle is that a reagent is a chemical which can be taken out of a bottle or container. Failure to identify whole reagents **will be penalised**. The command word “Identify”, allows the candidate to choose to use either the name or the formula in their answer. In some circumstances, the list principle may apply when both are used.

For example

potassium cyanide rather than cyanide ion **or** KCN rather than CN^-
sodium hydroxide rather than hydroxide ion **or** NaOH rather than OH^-

F. Marking calculations, such as those involving enthalpy changes

In general

- The sign for an enthalpy change will be assumed to be positive unless specifically shown to be negative.
- A correct answer alone will score **full marks** unless the necessity to show working is specifically required in the question.
- A correct numerical value with the **wrong sign** will score **only one mark**.

All other values **gain no credit** except

- Two marks can be awarded for correct chemistry with an arithmetic error.
- One mark can be awarded for a correct mathematical statement (or cycle) for the method.

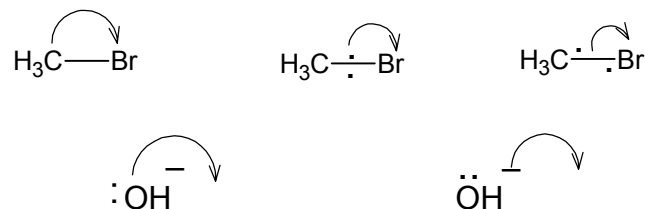
G. Oxidation states

In general, the sign for an oxidation state will be assumed to be positive unless specifically shown to be negative.

H. Organic reaction mechanisms

Curly arrows should originate either from a lone pair of electrons or from a bond.

Each of the following representations **should not gain credit** and will be penalised **once only** within a clip.



When the curly arrow is showing the formation of a bond to an atom, the arrow can go directly to the relevant atom, alongside the relevant atom or **more than half-way** towards the relevant atom.

In free-radical substitution

- The absence of a radical dot should be penalised **once only** within a clip.
- The use of double-headed arrows or the incorrect use of half-headed arrows in free-radical mechanisms should be penalised **once only** within a clip

I. Organic structures

In general

- Displayed formulae must show all of the bonds in the molecule, but need not show correct bond angles.
- Bonds should be drawn correctly between the relevant atoms.
For example, if candidates show the alcohol functional group as C-H-O, they should be penalised **on every occasion**.
- Some latitude should be given to the representation of C-C bonds in structures, given that CH₃— is considered to be interchangeable with H₃C— even though the latter would be preferred.
- Poor presentation of vertical C — CH₃ bonds or C — OH bonds or C — NH₂ bonds should **not** gain credit. The limit of tolerance is the half-way position between the vertical bond and the relevant atoms in the attached group.
- The use of 'sticks' in structures should **not** gain credit. The occasions that this applies will be indicated in the mark scheme.
- Some examples of formulae for specific compounds which should **not** gain credit are given here

CH₃COH for ethanalCH₃CH₂HO for ethanolOHCH₂CH₃ for ethanolC₂H₆O for ethanolCH₂CH₂ for etheneCH₂.CH₂ for etheneCH₂:CH₂ for ethene

N.B. Exceptions may be made in the context of balancing equations

Each of the following **should gain credit** as alternatives to correct representations of the structures.CH₂ = CH₂ for ethene, H₂C=CH₂CH₃CHOHCH₃ for propan-2-ol, CH₃CH(OH)CH₃

J. Organic names

As a general principle, non-IUPAC names or incorrect spelling or incomplete names should **not** gain credit.
Some illustrations are given here.

but-2-ol 2-hydroxybutane butane-2-ol 2-butanol	all should be butan-2-ol
2-methpropan-2-ol	should be 2-methylpropan-2-ol
2-methylbutan-3-ol	should be 3-methylbutan-2-ol
3-methylpentan 3-mythylpentane 3-methypentane	all should be 3-methylpentane
propanitrile	should be propanenitrile
aminethane	should be ethylamine (although aminoethane can gain credit)
2-methyl-3-bromobutane 3-bromo-2-methylbutane 3-methyl-2-bromobutane	all should be 2-bromo-3-methylbutane
2-methylbut-3-ene	should be 3-methylbut-1-ene
difluorodichloromethane	should be dichlorodifluoromethane